CHEMISTRY MOCK PAPER

【化學摘星模擬試卷】

- 完全針對HKDSE
- 附MC及LQ詳細題解
- 3 Sets Mock Papers
- 全面涵蓋考試題種及答題技巧

DSE
MOCK EXAM
CHEM

PAPER 1A

SET 1

CHEMISTRY PAPER 1

by Dr. Samuel Chong

2 hours 30 minutes

This paper must be answered in English

GENERAL INSTRUCTIONS

(1) There are **TWO** sections, A and B, in this Paper. You are advised to finish Section A in about 45 minutes.

(2) Section A consists of multiple-choice questions in this question paper, while Section B contains conventional questions printed separately in Question-Answer Book **B**.

(3) Answer to Section A should be marked on the Multiple-choice Answer Sheet while answers to Section B should be written in the spaces provided in Question-Answer Book **B. The Answer Sheet for Section A and the Question-Answer Book for Section B will be collected separately at the end of the examination.**

(4) A Periodic Table is printed on page 24 of Question-Answer Book **B**. Atomic numbers and relative atomic masses of elements can be obtained from the Periodic table.

INSTRUCTIONS FOR SECTION A (MULTIPLE-CHOICE QUESTIONS)

(1) Read carefully the instructions on the Answer Sheet. After the announcement of the start of the examination, you should first insert the information required in the spaces provided. No extra time will be given for filling information after the 'Time is up' announcement.

(2) When told to open this book, you should check that all the questions are there. Look for the words '**END OF SECTION A**' after the last question.

(3) All questions carry equal marks.

(4) **ANSWER ALL QUESTIONS.** You are advised to use an HB pencil to mark all the answers on the Answer sheet, so that wrong marks can be completely erased with a clean rubber. You must mark the answers clearly; otherwise you will lose marks if the answers cannot be captured.

(5) You should mark only **ONE** answer for each question. If you mark more than one answer, you will receive **NO MARKS** for that question.

(6) No marks will be deducted for wrong answers.

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Not to be taken away before the end of the examination session

This section consists of two parts. There are 24 questions in PART I and 12 questions in PART II.

Choose the best answer for each question.

Candidates may refer to the Periodic Table printed on page 24 of Question-Answer Book B.

PART I

1. Consider two elements A and B. It is known that they have the same number of outermost shell electron. Let *m* be the atomic number of A, given that the atomic number of B is larger than A, which of the following MUST NOT be the atomic number of B?

   A. $m + 2$
   B. $m + 6$
   C. $m + 8$
   D. $m + 18$

2. With reference to the following Periodic Table:

	I	II	III	IV	V	VI	VII	0
2		B		F		E		
3	A					C	D	

   (A to F are NOT the symbol of the elements)

   Given that all the elements above are in their standard form, which of the following statements is correct?

   A. A can react with oxygen to form a compound with pH smaller than 7.
   B. C is more reactive than E towards metal.
   C. During melting of F, numerous covalent bonds are broken.
   D. The compound formed between C and D conducts electricity at molten state.

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DSE

MOCK EXAM

CHEM

PAPER 1A

SET 2

HONG KONG DIPLOMA OF SECONDARY EDUCATION MOCK EXAMINATION

CHEMISTRY PAPER 1

(by Dr. Samuel Chong)

(2 hours 30 minutes)

This paper must be answered in English

GENERAL INSTRUCTIONS

(1) There are **TWO** sections, A and B, in this Paper. You are advised to finish Section A in about 45 minutes.

(2) Section A consists of multiple-choice questions in this question paper, while Section B contains conventional questions printed separately in Question-Answer Book **B**.

(3) Answer to Section A should be marked on the Multiple-choice Answer Sheet while answers to Section B should be written in the spaces provided in Question-Answer Book **B. The Answer Sheet for Section A and the Question-Answer Book for Section B will be collected separately at the end of the examination.**

(4) A Periodic Table is printed on page 26 of Question-Answer Book **B**. Atomic numbers and relative atomic masses of elements can be obtained from the Periodic table.

INSTRUCTIONS FOR SECTION A (MULTIPLE-CHOICE QUESTIONS)

(1) Read carefully the instructions on the Answer Sheet. After the announcement of the start of the examination, you should first insert the information required in the spaces provided. No extra time will be given for filling information after the 'Time is up' announcement.

(2) When told to open this book, you should check that all the questions are there. Look for the words '**END OF SECTION A**' after the last question.

(3) All questions carry equal marks.

(4) **ANSWER ALL QUESTIONS.** You are advised to use an HB pencil to mark all the answers on the Answer sheet, so that wrong marks can be completely erased with a clean rubber. You must mark the answers clearly; otherwise you will lose marks if the answers cannot be captured.

(5) You should mark only **ONE** answer for each question. If you mark more than one answer, you will receive **NO MARKS** for that question.

(6) No marks will be deducted for wrong answers.

Not to be taken away before the end of the examination session

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This section consists of two parts. There are 24 questions in PART I and 12 questions in PART II. Choose the best answer for each question.
Candidates may refer to the Periodic Table printed on the last page of the paper.

PART I

1. X and Y are two elements which form a compound with a formula of XY_2. Given that the melting point of X is 1750 ^oC and XY_2 is insoluble in any solvents. It is known that X in is Group IV. Which of the following does Y belong to?

   A. Group I
   B. Group II
   C. Group III
   D. Group VI

2. Which of the following statements is correct?

   A. Metallic bond is weak since metal layers can slide over one another when force is applied.
   B. Graphite has the lubricating effect due to weak van der Waals' forces between each layer.
   C. On melting phosphorus and potassium, numerous covalent bonds and metallic bonds are broken.
   D. The covalent bond between water molecules is so strong that molecules with simple molecular structures may not be easily dissolved in water.

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SET 3

CHEMISTRY PAPER I
(Section A - Version 1)

Time allowed: 2 hours 30 minutes

This paper must be answered in English.

GENERAL INSTRUCTIONS

1. There are **TWO** sections, A and B, in this Paper. Section A carries 36 marks and Section B carries 84 marks. You are advised to finish Section A in about 45 minutes and Section B in about 105 minutes.
2. Section A consists of multiple-choice questions in this question book, while Section B contains conventional questions printed separately in Question-Answer Book B.
3. Answers to Section A should be marked on the Multiple-choice Answer Sheet while answers to Section B should be written in the spaces provided in Question-Answer Book B.

SECTION A (MULTIPLE-CHOICE QUESTIONS)

INSTRUCTIONS FOR SECTION A

1. Read the instructions on the Answer Sheet carefully.
2. All questions carry equal marks.
3. **ANSWER ALL QUESTIONS**.
4. You should mark only **ONE** answer for each question. If you mark more than one answer, you will receive **NO MARKS** for that question.
5. No marks will be deducted for wrong answers.

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This section consists of two parts. There are 24 questions in **PART I** and 12 questions in **PART II**. Choose the best answer for each question.
Candidates may refer to the Periodic Table printed on the last page of the paper.

PART I

1. Which of the statement concerning metals is/are correct ?
   (1) Silver is expensive because it is inert.
   (2) Aluminium is the first found by human because it is very reactive.
   (3) Iron can be obtained by adding magnesium into iron(II) nitrate solution.

   A. (1) only
   B. (3) only
   C. (1) and (3) only
   D. (2) and (3) only

2. Which of the following compounds is/are trigonal planar shape ?
   (1) PCl_2F
   (2) NF_3
   (3) BCl_3

   A. (1) only
   B. (3) only
   C. (1) and (3) only
   D. (2) and (3) only

3. Consider the zinc-carbon cell, which of the following statement is correct ?

   A. The concentration of electrolyte increases during discharge.
   B. It is used as mobile phone battery.
   C. The negative electrode of the zinc-carbon cell is made of zinc.
   D. It is rechargeable.

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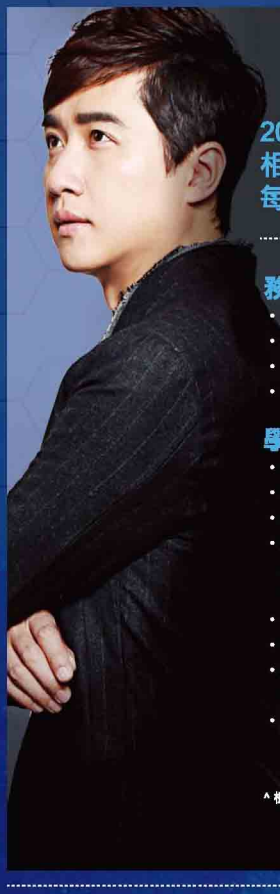
CHEMISTRY MOCK PAPER



2015 DSE化學科學生奪5摘星比率達58.2%^#
相當於全港同等成績的2.2倍
每1.7個學生就有1個取Level 5或以上

務實教學

- 英皇教育首席化學科名師，教材以高質素高效率見稱。
- 英皇教育博士級化學科名師。
- 香港大學化學系博士及榮譽學士畢業。
- 具有多年大學科研經驗，為化學科專家。

學生成績超卓

- 筆記均精美印刷，令學生清晰易明，大大提高課堂效率。
- 學生成績超卓，極多報讀學生奪取5**及A級成績^。
- 學生口碑絕讚，成績突飛猛進。
- 於04及05年應邀為學生福利聯盟（SWA）於沙田大會堂及灣仔會議展覽中心舉行之會考及高考講座作化學科主講嘉賓。於05至06年更應邀為元朗信義中學聯辦的會考講座中擔任化學科主講嘉賓。
- 於07-08、09-10年擔任英皇教育暑期講座之化學科主講嘉賓。
- 於09年擔任NOW TV「MOCK測高深」主講嘉賓。
- 於2015年擔任於浸會大學舉辦之「2016　DSE備戰策略講座」化學科唯一主講嘉賓。
- 具有多年教授補習經驗，對考試的趨勢瞭如指掌，並迅速地教授學生答題絕技。

^ 根據學生向英皇教育提供的數據或資料分析　#「奪5摘星」意指奪5級或以上成績

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